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alissaaauer. Terms in this set (32) Which

one of the following reactions is NOT

balanced?  $2 \text{CO} + \text{O}_2 \rightarrow 2 \text{CO}_2$   $2 \text{SO}_2 + \text{O}_2$

$\rightarrow 2 \text{SO}_3$   $\text{SF}_4 + 3 \text{H}_2\text{O} \rightarrow \text{H}_2\text{SO}_3 + 4 \text{HF}$   $2$

$\text{KNO}_3 + 10 \text{K} \rightarrow 5 \text{K}_2\text{O} + \text{N}_2$ .

Chapter 5 - Mastering Chemistry

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ANSWER: Correct Chapter 9 Reading

Question 2 Part A Iron reacts with oxygen

in the air to produce rust ( $\text{Fe}_2\text{O}_3$ )

according to the reaction shown below.  $4$

$\text{Fe} + 3 \text{O}_2 \rightarrow 2 \text{Fe}_2\text{O}_3$  How many moles of

rust can be formed if 1.8 mols of iron are

present? Hint 1. Begin with your given

number of moles.



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8 CHAPTER 8 REVIEW Chemical  
Equations and Reactions SECTION 3  
SHORT ANSWER Answer the following  
questions in the space provided. 1. List  
four metals that will not replace hydrogen  
in an acid. Choose from Cu, Ag, Au, Pt,  
Sb, Bi, and Hg.

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Mastering Chemistry Chapter 2: Rank from most to fewest valence electrons of the following elements: Kr, Ga, Br, As, K. - Answer: Kr (8), Br (7), As (5), Ga (3), K (1). Valence electrons are equal to the atom's main group number. Link: <https://www.coursehero.com/file/20333255/Mastering-Chemistry-Chapter-2/>.

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